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Chemistry Organic Chemistry: A Guided Inquiry Use curved arrows to show the most likely acid base reaction between each boxed pair. a. Use the pK_a values in Model 7 and in the previous ChemActivity to mark each curved arrow with a positive or negative energy change in pK_a Units. b. Calculate ΔH

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for each reaction, and sketch an energy diagram showing ΔH as an arrow on this diagram.

Use curved arrows to show the most likely acid base ...

Chemistry Organic Chemistry: A Guided Inquiry The equilibrium constant for the reaction from CTQ 47 is given below.

[Square brackets] represent the concentration of that ion or molecule in moles/liter at equilibrium. $K_{eq} = \frac{[NH_4^+] [OH^-]}{[NH_3] [H_2O]} = 10^{-7}$ a.

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